

The Crystal and Molecular Structure of Bis(thiocarbohydrazide-N,S)cadmium Dichloride

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The crystals of bis(thiocarbohydrazide-N,S)cadmium dichloride, $Cd(S=C(NH-NH_2)_2)_2Cl_2$, belong to monoclinic space group $P2_1/c$. The structure of the compound, determined from three-dimensional data, consists of trans-octahedral complexes where the ligand molecule forms pentatomic chelate rings with N,S as donor atoms. The chlorine atoms lie along the normal to the plane containing metal and organic molecules. The distances $Cd-N=2.34(3)$ and $Cd-S=2.59(1)$ Å are slightly shorter than those found in non-chelate cadmium complexes; $Cd-Cl=2.73(1)$ Å is very close to the sum of ionic radii. The configuration of the ligand, with one $-NH_2$ group turned towards $S=C$, repeats that of crystallized thiocarbohydrazide. Some differences, however, in distances, planarity, and angles between complexed or free ligand can be detected. The complexes are held together in layers parallel to (102) by weak hydrogen bonds $NH \dots Cl$.

Introduction

Thiocarbohydrazide, $S=C(NH-NH_2)_2$, forms complexes which can be assigned structures with pentatomic chelate rings with S,N as donor atoms. Such structures have been assumed to interpret IR spectra¹ and equilibria in solution.² In order to confirm these assumptions, the crystal structure of bis(thiocarbohydrazide-N,S)cadmium dichloride has now been determined.

Experimental Section

Preparation. Crystals of the compound, $Cd(SC-(NH-NH_2)_2)_2Cl_2$ have been obtained by evaporation of aqueous solutions of thiocarbohydrazide and cadmium chloride in stoichiometric ratio. Small, colorless crystals were obtained.

Crystal Data. Compound: bis(thiocarbohydrazide-N,S)cadmium dichloride, $Cd(SC(NH-NH_2)_2)_2Cl_2$, F.W. 395.62.

Crystal class: Monoclinic prismatic.

Unit cell (radiation $CuK\alpha$, $\lambda=1.5418$ Å, from

rotation and Weissenberg photographs around [100] and [010]):

$$a=8.64(1), \quad b=5.78(1), \quad c=13.78(1) \text{ Å};$$

$$\beta=119.5(3)^\circ, \quad V=599.3 \text{ Å}^3, \quad Z=2;$$

$$D_x=2.19, \quad D_m=2.17 \text{ g.cm}^{-3};$$

$$\mu(CuK\alpha)=224.7 \text{ cm}^{-1};$$

Space group $P2_1/c$ (C_{2h} (5), No. 14).

Intensity Data. Intensities of reflections $0kl, \dots, 7kl$, and $h0l, \dots, h5l$ were recorded on integrating Weissenberg camera and then measured by a microdensitometer (934 independent reflections out of 1369 possible).

Calculations. Usual corrections were applied. Absorption corrections as for cylindrical specimens were introduced ($\mu R_{[100]}=1.01$, $\mu R_{[010]}=0.67$). Atomic form factors according to Cromer and Mann³ were used.

Table I. Fractional atomic coordinates (with e.s.d.'s $\times 10^4$).

	x	y	z
Cd	1.0000	.0000	.0000
S	.6777(8)	-.1555(5)	-.0947(5)
Cl	.9028(8)	.2939(6)	-.1755(5)
N(1)	.6763(27)	.2649(17)	-.0094(18)
N(2)	.8578(34)	.2493(22)	.0655(19)
N(3)	.4141(24)	.1321(16)	-.1499(17)
N(4)	.3324(34)	.3398(26)	-.1450(26)
C	.5812(23)	.1021(24)	-.0864(17)

Table II. Anisotropic thermal parameters^a (Å²).

	B ₁₁	B ₂₂	B ₃₃	B ₁₂	B ₁₃	B ₂₃
Cd	2.825	3.130	3.248	0.131	1.197	0.042
S	2.951	2.898	3.361	-0.004	1.365	-0.139
Cl	3.067	3.076	3.153	-0.023	1.074	0.345
N(1)	2.561	3.378	3.249	0.109	0.657	-0.143
N(2)	3.663	3.046	3.767	-0.398	1.594	-0.788
N(3)	3.604	2.758	3.919	-0.100	1.550	0.283
N(4)	2.445	3.951	4.252	-0.026	0.977	0.128
C	2.800	3.166	2.987	-0.476	1.022	0.126

^a In the last cycle for all the atoms the average shift, $|\Delta B_{ij}|_{av}$, and the maximum shift, $|\Delta B_{ij}|_{max}$, were:

$$|\Delta B_{ij}|_{av} = 0.049$$

$$|\Delta B_{ij}|_{max} = 0.173$$

(1) G. R. Burns, *Inorg. Chem.* 7, 277 (1968).

(2) A. Braibanti, F. Dallavalle and E. Leporati, *Inorg. Chim. Acta*, 3, 459 (1969).

Table III. Observed and calculated structure factors= after F_o values indicates unobserved reflections.

h	k	l	$10F_O$	$10F_C$	h	k	l	$10F_O$	$10F_C$	h	k	l	$10F_O$	$10F_C$	h	k	l	$10F_O$	$10F_C$	h	k	l	$10F_O$	$10F_C$						
0	2	0	272	-239	1	1	1	390	403	1	0	8	289	329	2	1	-2	656	-711	2	3	-9	185	158	3	5	-3	220	201	
0	6	0	365	337	1	1	-1	668	655	1	0	8	80	-53	2	2	2	850	876	2	4	9	163	-123	3	6	3	111	99	
0	6	0	275	287	1	2	-1	458	-668	1	1	8	176	-150	2	2	-2	601	760	2	4	-9	159	-172	3	6	-3	80	60	
0	1	1	1056	1170	1	2	-1	155	-101	1	1	-8	187	-148	2	3	2	122	101	2	5	9	122	-126	3	7	-3	77	108	
0	2	1	325	-351	1	3	-1	759	831	1	2	-8	456	519	2	3	-2	102	63	2	5	-9	222	246	3	0	6	530	566	
0	3	1	371	440	1	3	-1	657	720	1	2	-8	525	619	2	6	2	259	242	2	6	-9	62	-39	3	0	-6	136	140	
0	4	1	301	332	1	4	-1	271	273	1	3	8	111	-16	2	6	-2	459	433	2	0	10	197	220	3	1	6	144	-119	
0	5	2	193	168	1	4	-1	97	98	1	3	-8	127	106	2	5	2	163	-126	2	0	10	161	172	3	1	4	144	-126	
0	6	1	100	-116	1	5	-1	361	336	1	4	8	108	-2	2	5	-2	100	-30	2	1	10	131	-107	3	2	6	210	168	
0	7	1	201	264	1	5	-1	272	214	1	4	-8	294	319	2	6	2	185	174	2	1	-10	159	-115	3	2	-4	904	917	
0	0	2	987	1082	1	6	1	99	-85	1	5	8	77	68	2	6	-2	191	183	2	2	10	207	157	3	3	4	101	-88	
0	1	2	265	341	1	6	-1	101	62	1	5	-8	121	-72	2	7	2	32	14	2	2	-10	427	475	3	3	-4	265	233	
0	2	2	537	630	1	7	1	63	53	1	6	8	175	189	2	7	-2	405	-7	2	3	10	52	-36	3	4	6	510	528	
0	3	2	115	-13	1	7	-1	133	204	1	6	-8	68	-67	2	1	3	797	819	2	3	-10	92	-17	3	6	-4	424	378	
0	4	2	345	409	1	0	2	163	105	1	1	9	385	403	2	1	-3	436	460	2	4	10	251	224	3	5	6	129	111	
0	5	3	111	67	1	0	2	848	895	1	1	-9	301	337	2	2	2	355	356	2	4	-10	166	171	3	5	-6	205	-183	
0	6	2	211	176	1	1	2	169	170	1	2	9	131	-29	2	2	3	203	-190	2	5	10	89	-57	3	6	4	97	70	
0	7	2	63	-19	1	1	-2	64	-35	1	2	-9	126	-65	2	3	3	240	215	2	6	-10	52	79	3	6	-4	77	51	
0	1	3	755	794	1	2	2	904	959	1	3	9	109	-126	2	3	-3	697	683	2	1	11	173	144	3	7	-4	41	30	
0	2	3	207	-208	1	2	-2	715	753	1	3	-9	283	304	2	4	3	99	263	2	1	-11	246	235	3	1	5	172	205	
0	3	4	340	386	1	3	2	232	-232	1	4	9	101	-13	2	4	-3	141	-101	2	2	11	48	19	3	1	5	594	657	
0	4	5	206	165	1	3	-2	169	117	1	4	-9	91	-8	2	5	3	284	285	2	2	-11	114	9	3	2	5	206	-161	
0	5	3	443	449	1	4	2	412	434	1	5	9	166	178	2	5	-3	371	377	2	3	11	284	239	3	2	-5	264	-244	
0	6	3	97	-35	1	4	-2	311	305	1	5	-9	235	272	2	6	3	100	95	2	3	-11	319	317	3	3	-5	564	548	
0	7	3	802	45	1	5	2	93	67	1	6	-9	55	-2	2	6	-3	186	179	2	4	11	60	-19	3	3	-5	473	448	
0	8	0	34	-62	1	5	-2	126	103	1	0	10	321	329	3	7	3	20	135	2	4	-11	82	-22	3	4	5	147	106	
0	1	6	786	-840	1	6	2	102	107	1	0	-10	181	137	2	7	-3	40	36	2	5	-11	104	162	3	4	5	91	84	
0	2	6	626	716	1	6	-2	240	269	1	1	10	178	-163	2	0	4	1283	1391	2	0	12	186	181	3	5	5	147	125	
0	3	4	219	236	1	7	2	56	11	1	1	-10	122	103	2	0	4	6	421	435	2	0	12	195	199	3	5	5	284	292
0	4	6	362	508	1	7	-2	82	-118	1	2	10	215	221	2	1	4	371	-393	2	1	12	131	127	3	6	5	45	-58	
0	5	5	100	72	1	1	3	513	623	1	2	-10	309	348	2	1	-6	634	-692	2	1	-12	100	71	3	6	3	76	-22	
0	6	6	92	-68	1	1	-3	1234	1327	1	3	10	102	-59	2	2	4	134	125	2	2	12	219	216	3	7	-5	36	27	
0	7	7	52	-48	1	2	3	171	170	1	3	-10	159	-108	2	2	-6	1101	1160	2	2	-12	237	195	3	0	6	185	198	
0	1	5	232	289	1	2	-3	70	68	1	4	10	116	160	2	3	-6	149	131	2	3	12	51	-31	3	0	-6	606	691	
0	2	6	50	94	1	3	-3	493	537	1	4	-10	249	260	2	3	-6	67	-18	2	3	12	82	-58	3	1	6	130	125	
0	3	5	745	849	1	3	-2	58	-10	1	5	10	51	-55	2	4	-6	250	238	2	6	-11	296	277	3	1	-6	80	-59	
0	4	6	116	-113	1	4	-3	197	-197	1	5	-10	83	56	2	4	-6	308	283	2	5	-12	76	77	3	6	2	300	365	
0	5	5	144	128	1	4	-3	72	-32	1	1	11	131	120	2	5	4	73	-73	2	1	-13	178	196	3	2	6	390	416	
0	6	5	87	-86	1	5	3	321	349	1	1	-11	215	226	2	5	-6	195	174	2	3	14	151	138	3	3	6	166	-155	
0	7	5	73	-122	1	5	-3	412	410	1	2	11	119	-60	2	6	4	210	229	2	2	13	41	-31	3	3	6	219	-176	
0	0	6	697	728	1	6	3	92	-66	1	2	-11	250	-226	2	6	-6	223	222	2	2	-13	105	-105	3	3	4	336	320	
0	1	6	70	-57	1	6	3	97	-61	1	3	11	291	304	2	7	-2	307	-37	2	3	13	304	-304	3	3	4	355	-401	
0	2	6	262	292	1	7	3	52	-29	1	3	11	289	287	2	6	-2	277	262	2	2	15	76	-69	3	3	4	355	-401	
0	3	7	133	155	1	7	4	69	199	1	2	12	198	219	2	5	5	321	311	2	3	15	60	-57	3	3	7	277	263	
0	4	8	7	100	-86	1	6	3	211	216	1	2	16	304	308	2	6	-6	383	436	3	0	6	268	295	3	1	-8	304	-348
0	5	9	131	128	1	1	-5	797	854	1	3	13	112	143	2	3	6	92	-36	3	7	0	30	36	3	2	8	233	219	
0	6	10	111	-84	1	2	-5	227	220	1	3	-13	131	129	2	3	-6	80	-64	1	1	1	835	869	3	2	-8	111	100	
0	7	11	118	-122	1	2	-5	98	-33	1	4	-13	65	-27	2	4	6	211	221	1	1	-1	637	654	3	3	8	133	97	
0	8	5	50	-67	1	6	3	92	-136	1	1	-15	186	197	2	1	-2	681	719	3	5	-1	367	443	3	6	7	232	221	
0	9	11	287	323	1	7	6	140	164	1	2	-15	72	-54	2	2	7	198	-176	3	6	1	129	135	3	3	-8	82	86	
0	10	2	11	129	1	0	6	696	727	1	0	-16	96	161	2	2	-7	163	118	3	6	1	106	-106	3	2	-9	307	70	
0	11	3	111	156	1	0	6	873	928	1	1	-16	36	151	2	3	2	225	212	2	7	1	222	227	3	3	9	227	166	
0	12	4	111	89	1	1	6	271	313	2	0	-6	555	563	2	3	-7													

Table III. (cont.)

h	k	l	$\log_{\frac{1}{2}}$	$\log_{\frac{1}{2}}$	h	k	l	$\log_{\frac{1}{2}}$	$\log_{\frac{1}{2}}$	h	k	l	$\log_{\frac{1}{2}}$	$\log_{\frac{1}{2}}$	h	k	l	$\log_{\frac{1}{2}}$	$\log_{\frac{1}{2}}$	h	k	l	$\log_{\frac{1}{2}}$	$\log_{\frac{1}{2}}$						
4	1	-6	493	549	5	4	1	243	-199	5	0	10	16-	83	6	3	-5	486	612	7	4	2	204	228	8	3	-1	244	229	
4	2	6	422	381	5	5	-1	163	133	5	0	-10	86-	129	6	4	5	51-	15	7	4	-2	158	113	8	4	-1	77	-21	
4	2	-6	473	515	5	5	-1	203	230	5	1	-10	37-	5	6	4	-5	112-	73	7	5	-2	65-	78	8	0	-2	89	66	
4	3	-6	70-	25	5	5	-1	208	197	5	2	-10	395	406	6	5	-5	156	153	7	1	3	169	113	8	0	-2	99	66	
4	3	-6	110	-82	5	5	1	64-	19	5	3	-10	136	132	6	6	-5	80	93	7	1	-3	351	349	8	1	-2	87-	103	
4	4	-6	129	86	5	6	-1	77	-57	5	4	-10	181	192	6	0	6	84-	48	7	2	3	51-	-38	8	1	-2	116	-39	
4	4	-6	409	350	5	0	2	409	429	5	3	-10	102	-78	6	0	-6	176	146	7	2	-3	163	-162	8	2	2	101-	100	
4	5	-6	65-	-43	5	6	-1	77	-57	5	4	-10	181	192	6	0	6	176	146	7	2	-3	311	292	8	2	-2	306	362	
4	5	-6	107	-65	5	1	-2	457	521	5	6	-10	26-	107	6	1	6	151-	-70	7	3	3	292	195	8	3	2	54-	-20	
4	6	-6	95	93	5	1	-2	80-	-74	5	2	-11	207	177	6	2	6	215	262	7	4	3	50	74	8	3	-2	95-	-29	
4	1	7	370	346	5	2	-2	361	324	5	3	-11	237	261	6	2	-6	508	513	7	6	-3	87-	78	8	4	-2	137	136	
4	1	-7	372	413	5	2	-2	617	379	5	6	-11	99-	-77	6	3	6	76	45	7	5	-3	95-	129	8	1	-3	169	183	
4	2	7	151	142	5	3	2	237	227	5	3	-11	190	186	6	3	-6	105-	66	7	0	-6	55-	16	8	1	-2	211	236	
4	3	7	86-	28	5	6	-2	204	190	5	1	-12	126	86	6	5	-6	90	80	7	1	-6	183	156	8	2	-3	166-	-112	
4	4	7	72-	-73	5	5	-2	149	165	5	3	-12	60-	-47	6	1	7	115	97	7	2	4	262	257	8	4	-3	90-	116	
4	5	-7	20-	202	5	6	-2	20-	157	5	5	-12	140	131	6	1	-7	333	354	7	2	-6	101-	85	8	5	-3	16-	101	
4	5	-7	126	109	5	6	-2	86	70	5	1	-13	200	218	6	2	-7	255	-247	7	3	-4	80-	50	8	0	-6	306	426	
4	6	-8	346	340	5	1	-3	296	280	5	3	-13	227	200	6	4	-7	165	148	7	5	-6	31-	30	8	1	-2	119-	-66	
4	0	-8	518	602	5	2	3	201	198	5	4	-13	126	104	6	5	-7	196	195	7	1	5	180	181	8	2	-2	151-	181	
4	1	8	150	-128	5	2	-3	195	-166	5	5	-13	28-	66	6	6	-7	51-	-23	7	1	-5	233	233	8	3	-6	102-	-59	
4	1	9	166	181	5	3	3	291	281	5	0	-6	271	252	6	0	8	158	219	7	2	5	52	68	8	4	-2	97-	98	
4	2	8	148	117	5	3	-3	742	680	5	1	-14	153	-93	6	0	-8	308	361	7	2	-5	266	257	8	5	-6	35-	68	
4	2	-8	249	262	5	4	3	92	-71	5	2	-14	187	137	6	1	8	45-	78	7	3	-5	357	316	8	1	-6	333	348	
4	3	8	61-	39	5	6	-3	90-	18	5	3	-16	83-	-30	6	1	-8	101-	-11	7	4	-5	161	-170	8	2	-3	151-	-66	
4	3	-8	96-	-42	5	5	-3	116	99	5	4	-16	126	116	6	2	-8	227	223	7	5	-5	112	140	8	3	-3	307	117	
4	4	8	126	124	5	5	-3	105	87	5	1	-15	163	154	6	3	-8	148	-127	7	0	-6	161	120	8	4	-5	97-	-28	
4	4	-8	457	625	5	6	-3	67	58	5	2	-15	101	-86	6	4	-8	211	179	7	0	-6	119	104	8	5	-5	37-	61	
4	5	-8	80-	-17	5	0	4	402	430	5	3	-15	143	132	6	5	-8	126	88	7	1	6	67-	17	8	0	-6	267	301	
4	6	-8	34-	22	5	0	-6	532	576	5	0	-16	155	160	6	6	-8	162	136	7	1	-6	118	98	8	1	-6	164	-190	
4	1	9	154	150	5	1	-6	121	113	5	1	-16	115	-127	6	1	-9	355	301	7	2	-6	399	405	8	2	-6	261	273	
4	1	-9	319	359	5	1	-6	589	-604	5	2	-16	163	136	6	2	-9	62-	46	7	3	-6	82	60	8	3	-6	107-	41	
4	2	9	87	-69	5	2	-6	211	206	5	3	-16	65-	-29	6	3	-9	203	204	7	4	-6	168	175	8	4	-6	97-	36	
4	2	-9	62-	-51	5	2	-6	392	381	5	1	-17	112	162	6	4	-9	80-	27	7	5	-6	90-	60	8	5	-6	40-	-10	
4	3	9	190	183	5	3	4	77	-65	5	2	-17	48-	-7	6	5	-9	76	63	7	1	-7	299	295	8	1	-2	223	346	
4	3	-9	451	428	5	3	-4	126	101	6	0	-6	305	421	6	0	-10	493	515	7	2	-7	132	125	8	2	-2	151-	1	
4	4	-9	88	-67	5	6	4	139	117	6	1	0	104	-101	6	1	-10	191	-191	7	3	-7	254	220	8	3	-7	107-	17	
4	5	-9	116	80	5	6	-4	323	298	6	2	0	287	276	6	2	-10	184	171	7	4	-7	61-	-23	8	4	-1	164	-190	
4	6	-9	126	84	5	5	-5	47-	10	6	3	0	131	96	6	3	-10	87	79	7	5	-7	173	203	8	5	-7	199	202	
4	0	10	233	221	5	5	-4	100	-77	6	4	6	0	195	162	6	4	-10	124	102	7	0	-8	77-	69	8	0	-6	107-	41
4	0	-10	213	209	5	6	-4	155	149	6	5	0	86	-68	6	5	-10	65-	-35	7	1	-8	183	184	8	1	-6	119-	165	
4	1	10	51-	7	5	1	-6	361	302	6	1	-1	341	347	6	1	-11	313	349	7	2	-8	350	375	8	2	-8	264	298	
4	1	-10	195	-184	5	1	-5	607	674	6	1	-1	393	389	6	2	-11	79	90	7	3	-8	115	-98	8	3	-8	102-	-23	
4	2	10	101	101	5	5	2	5	204	-162	6	2	-1	76-	-37	6	3	-11	107-	63	7	4	-8	247	228	8	4	-6	97-	94
4	3	10	373	427	5	2	-5	108-	-25	6	2	-1	203	183	6	4	-11	143	175	7	5	-6	108-	-108	8	3	-6	35-	157	
4	4	-10	155	-71	5	3	-5	116	90	6	3	-1	163	175	6	5	-11	144	183	7	1	-9	217	201	8	1	-10	116-	164	
4	5	-10	115	-83	5	3	-5	68-	62	6	3	-1	166	171	6	6	-12	103	103	7	2	-9	171	-140	8	2	-8	146-	122	
4	4	-10	191	192	5	4	5	109	-93	6	4	-1	87	-79	6	1	-12	97	66	7	3	-9	367	356	8	3	-8	264-	235	
4	5	-10	124	109	5	6	-5	157	-53	6	4	-1	208-	-24	6	5	-10	204	187	7	4	-10	195	204	8	4	-6	109-	155	
4	5	-11	212	208	5	1	-6	86	65	6	1	-2	348	355	6	4	-13	204	191	7	5	-6	155-	165	8	1	-11	112	117	
4	0	11	259	284	5	2	-6	308	310	6	2	-2	133-	70	6	0	-16	175	171	7	1	-11	195	233	8	2	-2	140-	95	
4	1	-12	107-	66	5	2	-6	522	519	6	2	-2	259	226	6	1	-16	105-	93	7	2	-11	67-	62	8	3	-11	255	260	
4	2	-12	267	226	5	3	-6	48-	-23	6	3	-2	105-	30	6	2	-16	147	113	7	3	-11	187	190	8	4	-11	80-	-70	
4	3	-12	191	-165	5	3	-6	133	-97	6	3	-2	60-	-9</																

Table IV. Atomic peak heights ($\text{e}.\text{\AA}^{-3}$), curvatures ($\text{e}.\text{\AA}^{-5}$), and e.s.d.'s.

	ρ	A_{hk}	A_{kk}	A_{ll}	A_{hk}	A_{kl}	A_{ll}
Cd	obs. calc.	109.1 110.5	1028 1030	1084 1087	1017 1032	12 9	460 470
S	obs. calc.	31.1 31.3	304 305	319 317	280 286	-3 -2	134 137
Cl	obs. calc.	31.2 31.6	278 282	297 298	283 286	-1 0	121 125
N(1)	obs. calc.	10.7 11.0	79 78	98 100	79 83	-2 -4	32 36
N(2)	obs. calc.	9.7 9.8	71 73	78 79	86 88	-7 -7	37 38
N(3)	obs. calc.	10.6 10.7	99 100	84 83	85 85	-3 -4	42 43
N(4)	obs. calc.	8.2 8.2	70 69	61 62	57 59	0 1	29 30
C	obs. calc.	9.9 10.1	113 115	71 73	85 85	-10 -7	45 48
e.s.d.'s		3	4	5	6	3	3

Table V. Main interatomic distances and angles (with e.s.d.'s).

Coordination around metal

Cd-S	2.59(1) Å	S-Cd-N(2)	76.7(6)°
Cd-Cl	2.73(1)	S-Cd-Cl	89.1(2)
Cd-N(2)	2.34(3)	Cl-Cd-N(2)	88.1(5)

Thiocarbohydrazide molecule

S-C	1.74(2)	Cd-S-C	95.8(7)
C-N(1)	1.35(2)	S-C-N(1)	121.5(17)
C-N(3)	1.28(3)	S-C-N(3)	119.2(13)
N(1)-N(2)	1.39(4)	C-N(1)-N(2)	124.3(17)
N(3)-N(4)	1.41(3)	C-N(3)-N(4)	120.4(18)
		N(1)-N(2)-Cd	111.5(14)
		N(1)-C-N(3)	119.2(17)

Hydrogen bonds

N(2)-Cl ⁱ	3.23(2)	Cd-N(2)-Cl ⁱ	110.3(11)
N(3)-Cl ⁱⁱ	3.26(2)	N(1)-N(2)-Cl ⁱ	121.4(9)
N(4)-N(1 ⁱⁱⁱ)	3.15(3)	C-N(3)-Cl ⁱⁱ	132.8(9)
		N(4)-N(3)-Cl ⁱⁱ	106.7(16)
		N(3)-N(4)-N(1 ⁱⁱⁱ)	146.0(20)
		N(4)-N(1 ⁱⁱⁱ)-N(2 ⁱⁱⁱ)	86.5(13)
		N(4)-N(1 ⁱⁱⁱ)-Cl ⁱⁱⁱ	145.9(19)

Asymmetric units

i $-x+2, -y+1, -z$ ii $-x+1, y-\frac{1}{2}, -z-\frac{1}{2}$ iii $-x+1, -y+1, -z$

The structure was solved by Patterson and Fourier methods and refined by differential syntheses.

Anisotropic temperature factors were introduced. (Final $R=9.7\%$). Hydrogen atoms were not identified in the difference Fourier map. The results of the structure determination are summarized in Tables I-V.

Discussion

The structure (Figure 1) consists of octahedral complexes with chelate molecules lying in the same plane and the chlorine atoms in the apical positions.

The chelating atoms (Figure 2) are N, of $-\text{NH}_2$, and S, thus confirming the assumptions on which were based the interpretations of IR spectra¹ and of

equilibria in solution.² Atoms N, S are in *trans* positions and all of them lie by symmetry conditions in the same plane as the metal atom. The chlorine atoms lie very close to the normal to that plane. The distance Cd-Cl = 2.73(1) Å is comparable with the sum of ionic radii (2.78 Å) and shows that the bond should be prevalently ionic. The organic ligand forms pentatomic chelate rings as expected; it is therefore in the same configuration, with one $-\text{NH}_2$ group turned toward S=C, as in the neutral molecule in crystals.⁴ This configuration is different from that found in thiocarbohydrazide dichloride⁵ where both $-\text{NH}_2$ groups are turned toward S=C.

(3) D. T. Cromer and J. B. Mann, *Acta Cryst.*, A24, 321 (1968).

(4) A. Braibanti, A. Tiripicchio and M. Tiripicchio Camellini, *Acta Cryst.*, B25, 2286 (1969).

(5) A. Braibanti, M. A. Pellinghelli, A. Tiripicchio and M. Tiripicchio Camellini, *Inorg. Chim. Acta* (in the press).

The distance Cd—N=2.34(3) Å is only slightly shorter than that found in chelates of hydrazinecarboxylic acid (Cd—N=2.38(3) Å,⁶ 2.40(2) Å⁷ and is equal to Cd—O in octahedral complexes (Cd—O=2.34(3) Å⁸). The distance Cd—S=2.59(1) Å is significantly shorter than that found in complexes of thiourea (Cd—S=2.638(4), 2.647(4) Å⁹); the shortening is probably due to chelation.

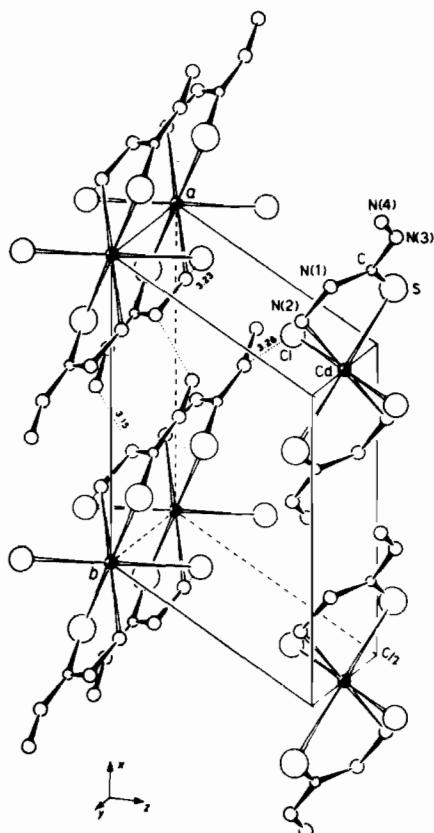


Figure 1. Clinographic projection of the structure of bis(thiocarbohydrazide-N,S)cadmium dichloride.

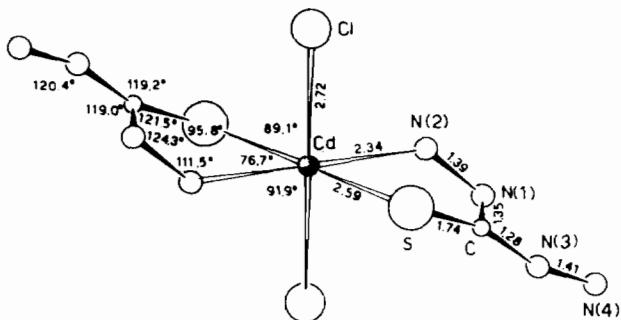


Figure 2. Trans-octahedral complex.

(6) A. Braibanti, A. Tiripicchio, A. M. Manotti Lanfredi and F. Bigoli, *Zeit. Kristallogr.*, **126**, 307 (1968).

(7) A. Braibanti, A. M. Manotti Lanfredi, A. Tiripicchio and F. Bigoli, *Acta Cryst.*, **B25**, 100 (1969).

In the chelate ligand some differences with the isolated molecule can be detected. The distance C—S=1.74(2) Å is comparable with that in thiocarbohydrazide (C—S=1.724(10) Å⁴), in complexed thiourea (C—S=1.76(1) Å,⁹ 1.73(1) Å¹⁰ and in free thiourea (C—S=1.71(1) Å¹¹). Also the distances N(1)—N(2)=1.39(4) Å and N(3)—N(4)=1.41(3) Å are not different from one another and equal to that found in the free ligand (1.405(8) Å⁴). In any case they are shorter than those found in free (N—N=1.46 Å¹²) and complexed hydrazine (N—N=1.47(2)¹³). The distances C—N(1)=1.35(2) and C—N(3)=1.28(3) Å as compared with single bond C—N=1.47 Å indicate that they possess some double bond character, although in different amounts. They can be compared with values found in thiocarbohydrazide (C—N=1.327(7) Å⁴), in thiourea (C—N=1.33(1) Å¹¹), and in thiosemicarbazide (C—N=1.323(2) Å¹⁴). The angle S—C—N(1)=121.5(17)° is smaller than the corresponding angle in the free molecule.

Table VI. Planarity of thiocarbohydrazide molecule. Equation of plane $5.2772x + 2.4606y - 12.1516z - 4.3449 = 0$

	Δ	σ
Cd	0.9323 Å	0.0000 Å
*S	-0.0004	0.0075
*C	0.0233	0.0247
*N(1)	-0.0099	0.0264
*N(3)	-0.0130	0.0245
*N(2)	-0.0006	0.0297
*N(4)	0.0073	0.0369
$\Sigma(\Delta/\sigma)^2 = 1.3542$		

* Atoms lying in the plane.

The whole ligand molecule lies in one plane (Table VI). This aspect is again different from that found in the free ligand where both terminal —NH₂ groups are out of the plane of the thioureide group, N—CS—N.

The complexes are held together in the crystal by hydrogen bonds N(3)—H...Cl=3.26(2) Å and N(2)—H...Cl=3.23(2) Å, thus forming layers parallel to (102). Weak contacts N(1)...N(4)=3.15(3) Å join the layers to one another.

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